**Unit 6 EOC Concepts**

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| 1. Empirical Formula   *Remember you are making a tutorial so that you can use this to remember how to calculate empirical formula and molecular formula for EOC exam in May 2017. Chapter 7 test should demonstrate how quickly you can forget how to do this!!* | * Define * Give example of a formula that would follow that definition and a formula that would not follow that definition * Show detailed step by step determination of empirical formula of a compound that is 80.0% carbon and 20.0% hydrogen * A compound has an empirical formula of CH2O and a molar mass of 90.0 g/mol. Determine the molecular formula. Explain how you got the correct formula. |