

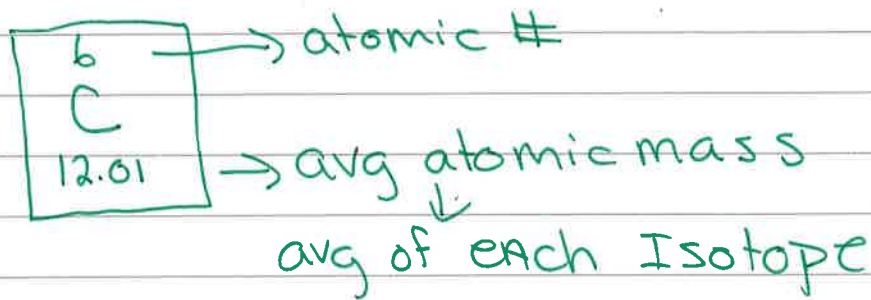
Unit 1 : Early Models of atom

Summary of models	Scientist	Name	Feature	Exp
	Dalton	Indivisible	Solid	non
	Thomson	Plum Pudding	e^-	Cathode Ray
	Rutherford	nuclear	p^+ in nucleus	Gold Foil
	Bohr	Planetary	e^- in set orbits	H-Line Emission

Sub atomic particles	Particle	Location	Charge	mass
	neutron	nucleus	none	1 amu
	proton	nucleus	+1	1 amu
	electron	outside nucleus	-1	0 amu

Atomic number & mass #

1) States the number of protons in an atom of an element



Isotope having same # protons, different # neutrons

Carbon-14

This represents one isotope of Carbon with a mass of 14 amu

Mass = $p + n$ so $14 - 6 = 8$
6 protons, 6 electrons, 8 neutrons